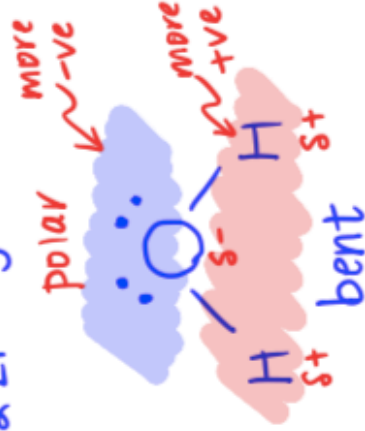


H<sub>2</sub>O



Around O:

2 bonds } 4 = bent  
2 LP



Around C:

Solution Chemistry Notes

4 bonds = tetrahedral  
no LP  
Non-polar



CO<sub>2</sub>

count  
1 double bond  
as 1 bond



Non Polar

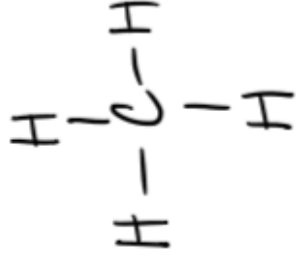
Dipole bonds cancel out

Around C: 2 bond  
NDLP

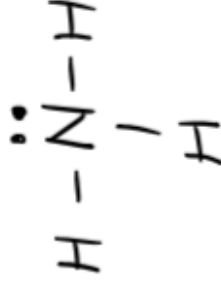
same

linear

CH<sub>4</sub>



NH<sub>3</sub>



Around N: 3 bonds } 4  
1 LP

= trigonal pyramid

